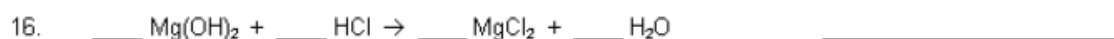
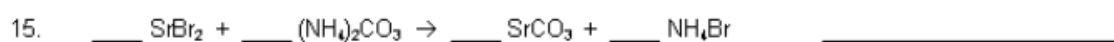
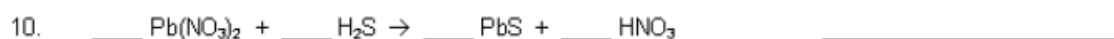
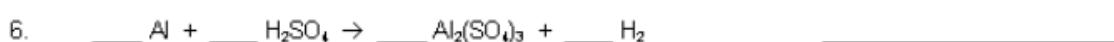
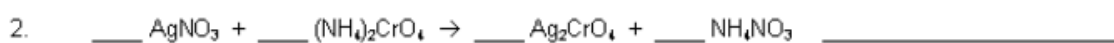


Balancing Practice and Answer Key



Balancing Practice and Answer Key

- $2 \text{ Sb} + 3 \text{ Cl}_2 \rightarrow 2 \text{ SbCl}_3$ synthesis
- $2 \text{ AgNO}_3 + 1 \text{ (NH}_4\text{)}_2\text{CrO}_4 \rightarrow 1 \text{ Ag}_2\text{CrO}_4 + 2 \text{ NH}_4\text{NO}_3$ double replacement
- $1 \text{ C}_3\text{H}_4 + 4 \text{ O}_2 \rightarrow 3 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ combustion
- $2 \text{ NaClO}_3 \rightarrow 2 \text{ NaCl} + 3 \text{ O}_2$ decomposition
- $1 \text{ Fe} + 2 \text{ HCl} \rightarrow 1 \text{ FeCl}_2 + 1 \text{ H}_2$ single replacement
- $2 \text{ Al} + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Al}_2(\text{SO}_4)_3 + 3 \text{ H}_2$ single replacement
- $2 \text{ CH}_2 + 3 \text{ O}_2 \rightarrow 2 \text{ CO}_2 + 2 \text{ H}_2\text{O}$ combustion
- $1 \text{ Na}_2\text{CO}_3 + 2 \text{ HCl} \rightarrow 2 \text{ NaCl} + 1 \text{ H}_2\text{CO}_3$ double replacement
- $2 \text{ C}_2\text{H}_6 + 7 \text{ O}_2 \rightarrow 4 \text{ CO}_2 + 6 \text{ H}_2\text{O}$ combustion
- $1 \text{ Pb}(\text{NO}_3)_2 + 1 \text{ H}_2\text{S} \rightarrow 1 \text{ PbS} + 2 \text{ HNO}_3$ double replacement
- $1 \text{ Zn} + 2 \text{ AgNO}_3 \rightarrow 1 \text{ Zn}(\text{NO}_3)_2 + 2 \text{ Ag}$ single replacement
- $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ decomposition
- $2 \text{ C}_4\text{H}_{10} + 13 \text{ O}_2 \rightarrow 8 \text{ CO}_2 + 10 \text{ H}_2\text{O}$ combustion
- $2 \text{ NaBr} + 1 \text{ Cl}_2 \rightarrow 2 \text{ NaCl} + 1 \text{ Br}_2$ single replacement
- $1 \text{ SrBr}_2 + 1 \text{ (NH}_4\text{)}_2\text{CO}_3 \rightarrow 1 \text{ SrCO}_3 + 2 \text{ NH}_4\text{Br}$ double replacement
- $1 \text{ Mg}(\text{OH})_2 + 2 \text{ HCl} \rightarrow 1 \text{ MgCl}_2 + 2 \text{ H}_2\text{O}$ double replacement
- $1 \text{ Ba}(\text{OH})_2 \rightarrow 1 \text{ BaO} + 1 \text{ H}_2\text{O}$ decomposition